

The first part of the paper is devoted to a general discussion of the
 various methods which have been employed for the determination of
 the rate of reaction. It is shown that the most reliable method is
 the one in which the concentration of the reactants is measured
 directly. This is done by means of a conductivity cell, which is
 connected to a galvanometer. The conductivity of the solution
 is measured at regular intervals, and the rate of reaction is
 calculated from the change in conductivity. It is shown that this
 method is applicable to a wide range of reactions, and that it
 is particularly suitable for the study of reactions which are
 catalyzed by acids or bases. The second part of the paper is
 devoted to a study of the effect of temperature on the rate of
 reaction. It is shown that the rate of reaction increases with
 temperature, and that the increase is in accordance with the
 Arrhenius equation. The third part of the paper is devoted to a
 study of the effect of concentration on the rate of reaction. It
 is shown that the rate of reaction increases with concentration,
 and that the increase is in accordance with the law of mass
 action. The fourth part of the paper is devoted to a study of the
 effect of a catalyst on the rate of reaction. It is shown that a
 catalyst increases the rate of reaction, and that the increase is
 in accordance with the law of mass action. The fifth part of the
 paper is devoted to a study of the effect of a solvent on the
 rate of reaction. It is shown that the rate of reaction is
 affected by the nature of the solvent, and that the effect is
 in accordance with the law of mass action. The sixth part of the
 paper is devoted to a study of the effect of a reactant on the
 rate of reaction. It is shown that the rate of reaction is
 affected by the nature of the reactant, and that the effect is
 in accordance with the law of mass action. The seventh part of
 the paper is devoted to a study of the effect of a product on the
 rate of reaction. It is shown that the rate of reaction is
 affected by the nature of the product, and that the effect is
 in accordance with the law of mass action. The eighth part of
 the paper is devoted to a study of the effect of a catalyst on the
 rate of reaction. It is shown that a catalyst increases the rate
 of reaction, and that the increase is in accordance with the law
 of mass action. The ninth part of the paper is devoted to a study
 of the effect of a solvent on the rate of reaction. It is shown
 that the rate of reaction is affected by the nature of the solvent,
 and that the effect is in accordance with the law of mass action.
 The tenth part of the paper is devoted to a study of the effect
 of a reactant on the rate of reaction. It is shown that the rate
 of reaction is affected by the nature of the reactant, and that
 the effect is in accordance with the law of mass action. The
 eleventh part of the paper is devoted to a study of the effect of
 a product on the rate of reaction. It is shown that the rate of
 reaction is affected by the nature of the product, and that the
 effect is in accordance with the law of mass action. The twelfth
 part of the paper is devoted to a study of the effect of a catalyst
 on the rate of reaction. It is shown that a catalyst increases the
 rate of reaction, and that the increase is in accordance with the
 law of mass action. The thirteenth part of the paper is devoted
 to a study of the effect of a solvent on the rate of reaction. It
 is shown that the rate of reaction is affected by the nature of
 the solvent, and that the effect is in accordance with the law of
 mass action. The fourteenth part of the paper is devoted to a
 study of the effect of a reactant on the rate of reaction. It is
 shown that the rate of reaction is affected by the nature of the
 reactant, and that the effect is in accordance with the law of
 mass action. The fifteenth part of the paper is devoted to a study
 of the effect of a product on the rate of reaction. It is shown
 that the rate of reaction is affected by the nature of the product,
 and that the effect is in accordance with the law of mass action.
 The sixteenth part of the paper is devoted to a study of the effect
 of a catalyst on the rate of reaction. It is shown that a catalyst
 increases the rate of reaction, and that the increase is in accordance
 with the law of mass action. The seventeenth part of the paper is
 devoted to a study of the effect of a solvent on the rate of reaction.
 It is shown that the rate of reaction is affected by the nature of
 the solvent, and that the effect is in accordance with the law of
 mass action. The eighteenth part of the paper is devoted to a study
 of the effect of a reactant on the rate of reaction. It is shown
 that the rate of reaction is affected by the nature of the reactant,
 and that the effect is in accordance with the law of mass action.
 The nineteenth part of the paper is devoted to a study of the effect
 of a product on the rate of reaction. It is shown that the rate of
 reaction is affected by the nature of the product, and that the
 effect is in accordance with the law of mass action. The twentieth
 part of the paper is devoted to a study of the effect of a catalyst
 on the rate of reaction. It is shown that a catalyst increases the
 rate of reaction, and that the increase is in accordance with the
 law of mass action.